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Mass Stability & Astrotheology

Paul T E Cusack, BScE, Dule

ABSTRACT

Simon's 12 th problem is the list of unsolved math-physics problems on Wikipedia states that we need to establish molecular structure from first principles. In this paper, that is what we do; from the individual atoms to the periodic table, to the benzene ring, we see the energy is always conserved. It is assumed that the reader is familiar with AT Math. If there was any doubt that Astro theology is the theory that explains the physical universe (as well as the spiritual), they can be laid to rest after reading this paper.

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Classification: LCC: QC170-197

Language: English



Great Britain
Journals Press

LJP Copyright ID: 925642
Print ISSN: 2631-8490
Online ISSN: 2631-8504

London Journal of Research in Science: Natural and Formal

Volume 23 | Issue 14 | Compilation 1.0



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I. INTRODUCTION

Simon's 12 th problem in the 2000 list:

Is there a mathematical sense in which one can justify current techniques for determining molecular configurations from first principles? Source: Wikipedia.

There are two forces in nature, one that draws together and the other that pushes apart. They are gravity and coulombic forces. We will show that the Mass in the universe is the result of these two opposing forces working against each other.

$$6.67/1.602=0.24017$$

$$\text{GMP} \Rightarrow t^2-t-1=E$$

$$0.2401^2-0.2401-1=1.18249=\text{Mass of the Period Table of the elements.}$$

The 118th element is Og.

$$118/118249=-0.99789$$

There are 32 elements in the 7th Period, or 32 elements. The Mass of an electron =5.1099mEv

$$-0.99789/(32 \times 5.1099)=6.10267$$

Orbitals

$$\{2+8+8+18+18+32+32\} \times 5.1099=602.9682=1/98.8 \approx 99$$

$$\text{Mass H}+=1.0079 \rightarrow \text{Mass Og}=294$$

$$\Delta M=292.99 \approx 293$$

$$292-392=198=t$$

$$198-981=693=\text{Ln } 2=M$$

$$693-936=297=c$$

$$297-792=495=E$$

$$495-549=99=E$$

$$99-99=0 \text{ Convergent}$$

$$198 \times 693 \times 297 \times 495 \times 99 = 1997 \sim 2$$

$$t=2$$

$$\text{GMP} \Rightarrow 2^2 - 2 - 1 = E = 1$$

Oganessian #118

$$118 \text{ p+ } \times 938.27208816$$

$$118 \text{ e- } \times 5.1099$$

$$118.4(943.3819) = 111696 + 1/2 = 1.617.46 \approx 1.618 \text{ Golden Mean}$$

GMP

$$1.618^2 - 1.618 - 1 = -76$$

$$1 - 0.76 = 0.24 \Rightarrow 1.602/6.67$$

Covalent radius for Og:

$$r = 157 = \pi/2$$

$$r = d = s = E \times t = |E| |t| \sin \theta$$

$$E = 1/\sin \theta = 1/F$$

$$\alpha = \omega^2 R = (d\theta/dt)R$$

$$\omega^2 = 157^2$$

$$= (157)^2 (293) (3/8)$$

$$= 271 \approx e^1 = E$$

Eigenvector

$$\sqrt{\sqrt{3}} = 1.316 = 1/0.759835 = 24.01$$

$$F = 2M\omega^2$$

$$= 2(292.99)(\pi/2)^2 = 1444$$

$$E = 1/F = 1/1444 = 692 = \text{Ln}1998 = M \text{ when } t=2$$

<table><tr><td>H</td><td colspan="15"></td></tr><tr><td>Hydrogen</td><td colspan="15"></td></tr><tr><td>1</td><td colspan="15"></td></tr><tr><td>1.008</td><td colspan="15"></td></tr></table>																H																Hydrogen																1																1.008																<table><tr><td>Al</td><td colspan="15"></td></tr><tr><td>Aluminum</td><td colspan="15"></td></tr><tr><td>13</td><td colspan="15"></td></tr><tr><td>26.982</td><td colspan="15"></td></tr></table>																Al																Aluminum																13																26.982																<table><tr><td>He</td><td colspan="15"></td></tr><tr><td>Helium</td><td colspan="15"></td></tr><tr><td>2</td><td colspan="15"></td></tr><tr><td>4.003</td><td colspan="15"></td></tr></table>																He																Helium																2																4.003																																																																																																																																																																																																																											
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85.468	87.62	88.906	91.224	92.906	95.95	98.907	101.07	102.906	106.42	107.868	112.414	114.818	118.710	127.4	127.6	126.905	131.294																																																																																																																																																																																																																																																																																																																																																																																																																																										
Cs	Ba	La		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn																																																																																																																																																																																																																																																																																																																																																																																																																																									
Cesium	Barium	Lanthanides		Hafnium	Tantalum	Tungsten	Rhenium	Osmium	Iridium	Platinum	Gold	Mercury	Thallium	Lead	Bismuth	Polonium	Astatine	Radon																																																																																																																																																																																																																																																																																																																																																																																																																																									
55	56	57-71		72	73	74	75	76	77	78	79	80	81	82	83	84	85	86																																																																																																																																																																																																																																																																																																																																																																																																																																									
132.905	137.327			178.49	180.948	183.85	186.207	190.23	192.22	195.08	196.967	200.59	204.38	207.2	208.980	209	210	222.018																																																																																																																																																																																																																																																																																																																																																																																																																																									
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Francium	Radium	Actinides		Rutherfordium	Dubnium	Seaborgium	Berkelium	Hassium	Mitnherium	Darmstadtium	Roganium	Copernicium	Nihonium	Flerovium	Moscovium	Livermorium	Tennessium	Oganesson																																																																																																																																																																																																																																																																																																																																																																																																																																									
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223.02	226.025			261	262	263	264	265	266	267	268	269	270	271	272	273	274	276																																																																																																																																																																																																																																																																																																																																																																																																																																									
<table><tr><td>La</td><td>Ce</td><td>Pr</td><td>Nd</td><td>Pm</td><td>Sm</td><td>Eu</td><td>Gd</td><td>Tb</td><td>Dy</td><td>Ho</td><td>Er</td><td>Tm</td><td>Yb</td><td>Lu</td></tr><tr><td>Lanthanum</td><td>Cerium</td><td>Praseodymium</td><td>Neodymium</td><td>Promethium</td><td>Samarium</td><td>Europium</td><td>Gadolinium</td><td>Terbium</td><td>Dysprosium</td><td>Holmium</td><td>Erbium</td><td>Thulium</td><td>Ytterbium</td><td>Lutetium</td></tr><tr><td>57</td><td>58</td><td>59</td><td>60</td><td>61</td><td>62</td><td>63</td><td>64</td><td>65</td><td>66</td><td>67</td><td>68</td><td>69</td><td>70</td><td>71</td></tr><tr><td>138.905</td><td>140.116</td><td>140.908</td><td>144.242</td><td>144.913</td><td>150.36</td><td>151.964</td><td>157.25</td><td>158.925</td><td>162.500</td><td>164.930</td><td>167.259</td><td>168.934</td><td>173.055</td><td>174.967</td></tr></table>																La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Lanthanum	Cerium	Praseodymium	Neodymium	Promethium	Samarium	Europium	Gadolinium	Terbium	Dysprosium	Holmium	Erbium	Thulium	Ytterbium	Lutetium	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	138.905	140.116	140.908	144.242	144.913	150.36	151.964	157.25	158.925	162.500	164.930	167.259	168.934	173.055	174.967	<table><tr><td>Ac</td><td>Th</td><td>Pa</td><td>U</td><td>Np</td><td>Pu</td><td>Am</td><td>Cm</td><td>Bk</td><td>Cf</td><td>Es</td><td>Fm</td><td>Md</td><td>No</td><td>Lr</td></tr><tr><td>Actinium</td><td>Thorium</td><td>Protactinium</td><td>Uranium</td><td>Neptunium</td><td>Plutonium</td><td>Americium</td><td>Curium</td><td>Berkelium</td><td>Californium</td><td>Einsteinium</td><td>Fermium</td><td>Mendelevium</td><td>Nobelium</td><td>Lr</td></tr><tr><td>89</td><td>90</td><td>91</td><td>92</td><td>93</td><td>94</td><td>95</td><td>96</td><td>97</td><td>98</td><td>99</td><td>100</td><td>101</td><td>102</td><td>103</td></tr><tr><td>227.028</td><td>232.038</td><td>231.036</td><td>238.029</td><td>237.048</td><td>244.064</td><td>243.061</td><td>247.065</td><td>247.065</td><td>251.080</td><td>252.083</td><td>257.105</td><td>258.10</td><td>259.108</td><td>262.105</td></tr></table>																Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	Actinium	Thorium	Protactinium	Uranium	Neptunium	Plutonium	Americium	Curium	Berkelium	Californium	Einsteinium	Fermium	Mendelevium	Nobelium	Lr	89	90	91	92	93	94	95	96	97	98	99	100	101	102	103	227.028	232.038	231.036	238.029	237.048	244.064	243.061	247.065	247.065	251.080	252.083	257.105	258.10	259.108	262.105																																																																																																																																																																																																																																																																																																				
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Figure 1: Periodic Table of the elements showing Og in lower right corner

We now use Benzene, a prolific molecule in organic chemistry to show why the atoms come together to conserve energy.

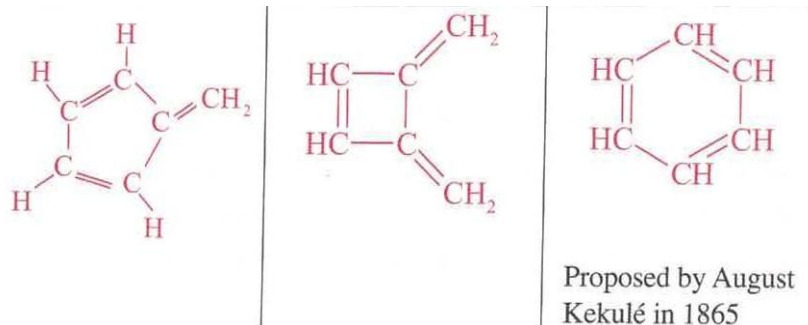


Figure 2: Benzene proposed structures. Source Barron's E Z Chemistry

$$3 \text{ C-C } 3(710)=2130$$

$$3 \text{ C-C } 3(607)=1821$$

$$6 \text{ C-H } 3(337.2)=2023.2$$

$$\text{SUM}=59742$$

$$\text{PE}+\text{Mc}^2=$$

$$\text{M}=6(12)+6(1)=78$$

$$\text{PE}=78(2.9979)^2=4222.3$$

$$\text{TE}=\text{PE}+\text{BDE}=59742+4222.3=10196$$

But the molecules all have the same number of each bond type.

Caron needs 4 e-

$$28 \times 1.602 = 4485.6$$

$$M = \ln t = 4485.6$$

$$T = 1.50 = 1/G$$

$$15 \text{ e- shared} \times 1.602 = 24.03 = 1.602 \times 6.67$$

So what is the difference for the Benzene ring? It is that 3 electrons are shared between 6 Carbons. So, energy is minimized:

$$3e-/6 \times (1.602) + 6(1.602) + He - 6(1.602) = 20.025$$

$$e^{20.025} = 2.99698 = c = 2.997 \sim 3$$

$t = c = 3 = \text{eigen value}$ $\square E = 5$ $\square y = y'$ for the eigen function.

For the middle architecture:

$$3 \text{ e-} + 6e- + 6e- = 20.025 \text{ same}$$

They are both symmetric molecules. Space must be conserved which is based on the cross product of energy and time(KE).

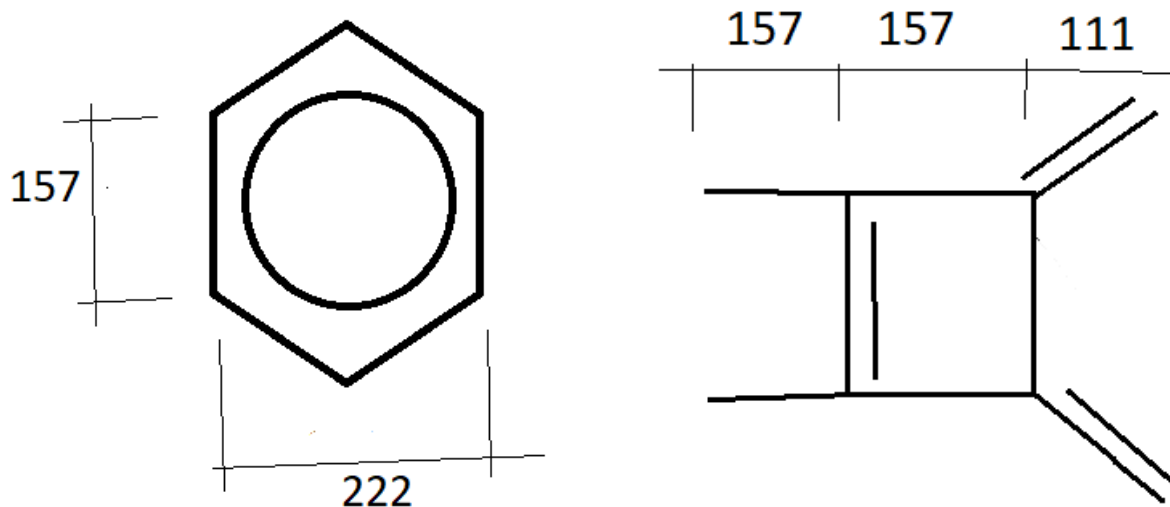


Figure 3: Two architectures for C₆H₆. The circle represents 3 electrons shared

$$s = \text{Ext} = E t \sin \theta$$

$$s = t$$

$$E = 1/\sin \theta = 20.025$$

$$\theta = 28.6 \text{ deg} = 0.5 \text{ rads}$$

$$\text{GMP: } 0.5^2 - 0.5 - 1 = -1.25 = E_{\text{min of the GMP}}$$

Area of the Benzene Ring:

$$S=(157)^2+(\pi/4)^2(\pi/4)^2/1/2$$

$$=375=1/F=E \text{ where } F=SF=8/3$$

Area of the middle structure:

$$157+157+111=425$$

$$111+157+111=379$$

Area -804

804

375 by 46.6% So the Benzene structure minimizes space. So that is why the Benzene ring is selected from first principles using AT Math.

II. CONCLUSION

The Mass of the universe are established from first principles of Astro theology. The joining up of atoms of elements follows the laws of the conservation of energy and space using coulombic forces.

ACKNOWLEDGEMENTS

I'd like to acknowledge Dr Owen Dunn of St Malachy's High School who taught me chemistry in 10 months. And Prof Carl Thompson at UNB Saint John who taught me Organic Chemistry for Health Sciences over a summer course. I also thank the government of New Brunswick for paying for my primary and secondary education, and the people of Canada for paying me a pension while I work at research. Finally, I thank my parents for educating me at home and allowing and encouraging me to learn while they worked to support us. Above all, I thank God for letting me push His pen; yet the many mistakes are all mine.

REFERENCES

1. Hathaway., BA., *Barron's E-Z Chemistry*. Barron's 2011.

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